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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle. $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$ $3 O_2 + 2 CO_2(g) + 3 H_2O - 1411 kJ mol^{-1} - 286 kJ mol^{-1} - 1560 kJ mol^{-1}$ (hydrogenation) (i) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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8202222 o correct reactants and products; [2] correct coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO / (unburnt) hydrocarbons; Do not accept methane/CH₄, ethane/C₂H₆, propane/C₃H₈ or butane/C₄H₁₀. Global pollutant: nitrogen oxide/NO / carbon ...

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$CH_3OH(l) + 3O_2(g) \rightarrow 2CO_2(g) + 3H_2O(l)$ $\Delta H = -1367 kJ mol^{-1}$ $CH_4(l) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(l)$ $\Delta H = -5470 kJ mol^{-1}$ (i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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2209-6110 - 3 - Turn over M09/4/CHEM/SPM/ENG/TZ1/XX+ 1. The molar mass of a compound is approximately 56 g mol⁻¹.Which formula is possible for this

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possible liver damage; [2 max] (b) mild analgesics function by intercepting the pain stimulus at the source / interfere with the production of substances that cause pain/prostaglandins; strong analgesics work by bonding to receptor sites in the brain / prevent the transmission of pain impulses without depressing the central nervous system; [2]

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