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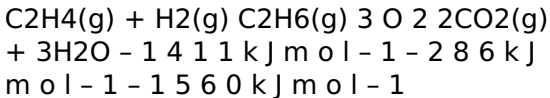
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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.



(hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of

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ethene from the energy cycle.....

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8 20 2 2 2 2 o correct reactants and products; [2] correct coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO /



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(unburnt) hydrocarbons; Do not accept methane/ $\text{CH}_4$ , ethane/ $\text{C}_2\text{H}_6$ , propane/ $\text{C}_3\text{H}_8$  or butane/ $\text{C}_4\text{H}_{10}$ .  
Global pollutant: nitrogen oxide/ $\text{NO}$  / carbon ...

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$\text{CH}_3\text{CHOH(l)} + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O(l)}$   $\Delta H = -1367 \text{ kJ mol}^{-1}$   
 $\text{C}_8\text{H}_{18}(\text{l}) + 12.5\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 9\text{H}_2\text{O(l)}$   $\Delta H = -5470 \text{ kJ mol}^{-1}$   
(i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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$(\text{CH}_3)_2\text{CHCH}_2\text{OH}$ ;  $(\text{CH}_3)_3\text{COH}$ ; [2 max]

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