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Write a balanced equation describing
each of the following chemical reactions.

(a) Solid potassium chlorate, KClO_3 ,

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decomposes to form solid potassium chloride and diatomic oxygen gas. (b) Solid aluminum metal reacts with solid diatomic iodine to form solid Al_2I_6 . (c) When solid sodium chloride is added to aqueous sulfuric acid, hydrogen chloride gas and aqueous sodium sulfate are produced.

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1. Transition elements can form coloured compounds.
2. Transition elements have varying oxidation number. Iron has two common oxidation states (+2 and +3) in, for example, Fe²⁺ and Fe³⁺.

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