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Name the three sub-atomic particles of an atom. Answer. The sub-atomic particles of an atom are- Electron, Proton and Neutron. 2.

Helium atom has an atomic mass of $4u$ and two protons in its nucleus. How many neutrons does it have? Answer. Atomic mass of Helium = $4u$

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Structure ...

Pop Quiz #4 The Atom

1. Who was the first to use the word atom? 2.

Dalton is known for developing the first theory of _____

structure. 3. What experiment helped

disprove the plum pudding model and showed the atom to have a nucleus in the

center. 4. Two particles found in the center of the atom are _____ and

_____ ^{5.}
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Instead of grams, the unit we use is the Atomic Mass Unit (amu) It is defined as one-twelfth the mass of a carbon-12 atom. Carbon-12 chosen because of its isotope purity. Each isotope

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has its own atomic mass, thus we determine the average from percent abundance.

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Chapter 4 Atomic
Structure Worksheet
Answer The Following
Questions Circle The
Best Answer 1 Which
Of The Following Are
Isotopes Of Each Other
A ^{14}C And ^{14}N B ^3H
And ^4He C ^2H And ^1H D
None Of These 2 The
Net Charge On An
Atom That Has 13
Protons 12 Neutrons
And 10 Electrons Is A 5
B 3 C 3 D 9 3 Particles
That Make Up Atoms
Are Called Subatomic

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Particles And Are

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...

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Structure - Madison
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4.2 The Structure of an
Atom Protons,
electrons, and
neutrons are
subatomic particles.

- A proton is a positively charged subatomic particle that is found in

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the nucleus of an atom.
It has a charge of $1+$.

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important themes
pervade science,
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at an ancient civilization, the human body, or a comet
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Science Chapter 4 –
Structure of The Atom
are provided here to
help students prepare
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Science exam. MCQs
from CBSE Class 9
Science Chapter 4:
Structure of The Atom
1. Who discovered the
electron?

MCQ Questions for

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Class 9 Science
Chapter 4 Structure
of ...

Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons, protons, and neutrons are examples of these fundamental particles There are many other types of particles, but we will study these three

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Atomic Structure - 4

...

Chapter 4 Assessment
Multiple Choice Identify
the choice that best
completes the
statement or answers
the question. ____ 1.

One of the first people
to state that matter is
made up of atoms was
a. Democritus. c.

Dalton. b. Aristotle. d.
Rutherford. ____ 2.

Dalton's model of an
atom is best described
as a. a solar system. c.

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a plum pudding. b ...

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Assessment - somer
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Chapter 4: Atomic
Structure Worksheet
Answer the following
questions, circle the
best answer 1) Which
of the following are
isotopes of each other?
a) ^{14}C and ^{14}N b) ^3H
and ^4He c) ^2H and ^1H
d) none of these 2) The
net charge on an atom
that has 13 protons,

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**[PDF] Chapter 4
Atomic Structure
Worksheet Answers**

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Chapter 4 - Atomic
Structure - 4.3
Distinguishing Among
Atoms - Sample
Problem 4.1 - Page 113
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by step written by
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Edition) Chapter 4 -
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...

Atomic charge =
number of protons –
number of electrons.

As will be discussed in
more detail later in this
chapter, atoms (and
molecules) typically
acquire charge by
gaining or losing

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electrons. An atom that gains one or more electrons will exhibit a negative charge and is called an anion.

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