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SECTION 10.1 THE

MOLE: A

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**MEASUREMENT OF
MATTER** (pages
287-296) This section
defines the mole and
explains how the mole
is used to measure
matter. It also teaches
you how to calculate
the mass of a mole of
any substance.

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SECTION 10.1 THE

MOLE: A

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287–296) This section

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any substance.

Measuring Matter

(pages 287–289)

Chapter 10 Chemical

Page 6/26

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Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. $1 \text{ mole} = 6.02 \times 10^{23}$ representative particles

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CHAPTER 10: Chemical Quantities BASICS: •

The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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CHAPTER 10:

Chemical Quantities

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Chapter 10 - Chemical
Quantities - 10.1 The
Mole: A Measurement
of Matter - 10.1 Lesson

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Check - Page 315: 9
Answer Knowing how
the count mass and
volume of an item
relate a common unit
allows you convert
among these units.

Chapter 10 Chemical Quantities Guided Reading Answer Key

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going in the same way
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91. SECTION
10.1 THE MOLE: A
MEASUREMENT OF
MATTER (pages

287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

Measuring Matter
(pages 287–289)

**SECTION 10.1 THE
MOLE: A**

Where To Download Chapter

10 Chemical **MEASUREMENT OF MATTER (pages 287-296)**

Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole.

atom 6.02×10^{23}

ion Na^+ 6.02×10^{23}

formula unit NaCl 6.02×10^{23}

representative particles of a

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**Chemical Quantities
- AP Biology**

10 1 kg 103 g 368

Chapter 10 Chemical
Calculations and
Chemical Equations.

10.1 Equation

Stoichiometry 369 The
ratio of moles of P 4O
10 to moles of P (which
came from the
subscripts in the
chemical formula, P 4O
10) provided the key

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conversion factor that
allowed us to convert

Answers

Chapter 10 Chemical Calculations and equations

Chapter 10 Chemical
Quantities. mole.

Avogadro's number.

representative particle.

molar mass. the SI unit

for measuring the

amount of a substance.

number of

representative

particles in a mole,

6.02×10^{23} . refers to

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the species present in
a substance: usually
atoms, m.... the mass
of one mole of a pure
substance.

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Chapter 10 - Chemical
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- The Mole: A

Measurement of Matter

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Quantities Packet (0)2

... Chapter 10:

Chemical Quantities. 1.

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Avogadro's Number or
NA. 2. empirical
formula. 3. molar
mass. 4. molar volume.
5 ...

Chapter 2: Physical Science Methods

Chapter 9 -
Stoichiometry ... Use
proportional reasoning
to determine mole
ratios from a balanced
chemical equation; ...
Section 9.1 Activities.
Section 9.1 Notes.
Chapter 9 Notes.

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Stoichiometry
Introduction Day;
Calculating Quantities
in Reactions Packet.
Calculating Quantities
in Reactions Packet.
Section 9.2 - Limiting
Reactants &
Percentage ...

Chapter 9 - Stoichiometry - Ms. Clark's Website

Chemistry Chapters 10
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Chapter 10: Chemical
Quantities, Chapter

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11: Chemical
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