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1  $Q = CV = 470 \times 10^{-6} \times 9.0 = 4.2 \times 10^{-3} \text{ C}$  [1] 2  $V = Q/C = 0.033 \text{ V}$  2200  $\times 10^{-6} = 15 \text{ V}$  [1] 3  $C = Q/V = 2.0 \text{ C} / 5000 = 4.0 \times 10^{-4} = 400 \mu\text{F}$  [1] 4  $W = 1/2 CV^2 = 0.5 \times 470 \times 10^{-6} \times 12^2 = 0.034 \text{ J}$

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$$[1] \quad 5 \text{ W} = 1 \quad 2 \text{ QV} = 0.5 \times 1.5 \times 10^{-3} \times 50 = 0.0375 \text{ J} \quad [1] \quad 6$$

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total mass of the nucleons ( $143m_n +$

$92m_p$ ). D Determine the mass defect  $\Delta m$

by subtracting the mass of the nucleus

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from the total mass of the nucleons. C  
Calculate the binding energy  $\Delta E$  of the  
nucleus using  $\Delta E = \Delta mc^2$ .

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62.1 = 5.18 1.0 10.3 = 10.3 16.0 27.6 =  
1.73 [1] Relative number of atoms 1.73  
5.18 = 3 1.73 10.3 = 6 1.73 1.73 = 1. 10  
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bC<sub>n</sub>H<sub>2n+1</sub>OH [1] cButan-1-ol and Page  
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